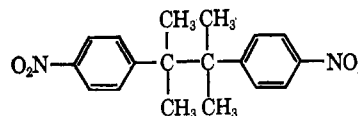


stable than the anion of malonic ester itself. On the other hand, the free radical A^{\cdot} (eq 6) will be more stable when derived from the monoalkylated malonic ester anion. The two effects reinforce each other and lead to a relatively high reactivity for the monoalkylated malonic ester anion.

Consistent with the proposed radical-anion mechanism is the fact that various of these reactions exhibit

some of the following characteristics: light speeds up the reaction, *p*-dinitrobenzene (or oxygen) retards the reaction, the dimer XII is isolated in 1-5% yields.⁷



XII

Aside from their mechanistic significance, the reactions of eq 1 have considerable interest because they provide a facile means of synthesizing highly ramified structures.

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(7) All reactions were conducted under nitrogen or argon.

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Additions and Corrections

Editor's Note: Hereafter additions and corrections to articles published in the *Journal of the American Chemical Society* will appear in this section promptly after they are received by the Editor. Authors are urged to submit such items when they become aware of the necessity, rather than deferring them to the final issue of the year.

Transition Metal Complexes of Triethylenetetramine. III. *cis*- α -Diacidotriethylenetetraminechromium(III) Complexes [*J. Am. Chem. Soc.*, **88**, 2156 (1966)]. By D. A. HOUSE and CLIFFORD S. GARNER. Department of Chemistry, University of California, Los Angeles, California 90024.

In Table I, the value 2950 in column 3 should be 2940, the value 1212 in column 4 should be 1312, and the value 1018 in column 7 should be 1028. On page 2161, in paragraph 2 of column 2, the calculated analysis of 32.9 is for Cl, not C.

Kinetics and Mechanism of the Gas Phase Reaction between Iodine and Isopropyl Alcohol and the Tertiary Carbon-Hydrogen Bond Strength in Isopropyl Alcohol [*J. Am. Chem. Soc.*, **88**, 3480 (1966)]. By R. WALSH and S. W. BENSON. Department of Thermochemistry and Chemical Kinetics, Stanford Research Institute, Menlo Park, California.

The following lines were inadvertently omitted from page 3481. They belong immediately after the last line of text on that page and before the first line of text on page 3482.

"...ent effect and was finally traced to losses (~1%) of *i*-C₃H₇OH to the grease. Since these could not be avoided, it was not possible to measure precisely

the importance of reaction D (also a pressure increase reaction). Analysis showed it to be almost negligible until beyond 50% reaction, and this was confirmed by"

Polarographic Study of Coordination Compounds with Delocalized Ground States. Substituent Effects in Bis- and Trisdithiodiketone Complexes of Transition Metals [*J. Am. Chem. Soc.*, **88**, 4876 (1966)]. By D. C. OLSON, V. P. MAYWEG, and G. N. SCHRAUZER. Shell Development Company, Emeryville, California, and The Institute of Inorganic Chemistry, The University, Munich, Germany.

The values of *m* in Table IV should have the opposite sign, *m* = 0 - 1, -2. On page 4882, column 2, line 5, ReS₆C₆Ph₆⁻² should read ReS₆C₆Ph₆⁻³.

The Synthesis of a 4',5'-Unsaturated Nucleoside [*J. Am. Chem. Soc.*, **88**, 5684 (1966)]. By J. P. H. VERHEYDEN and J. G. MOFFATT. Institute of Molecular Biology, Syntex Research, Palo Alto, California.

On page 5685, column 1, line 12, the formula C₁₃H₁₅N₂O₇ should read C₁₃H₁₅N₂O₇I. On the same page, column 1, line 26, the formula C₁₃H₁₄N₂O₆ should read C₁₃H₁₄N₂O₇.